

# Binary ionic compounds

Complete the following table:

<b>name</b>	<b>formula</b>
<b>lithium bromide</b>	
	<b>NaF</b>
<b>sodium oxide</b>	
	<b>Li<sub>2</sub>S</b>
<b>potassium phosphide</b>	
	<b>CaI<sub>2</sub></b>
<b>magnesium bromide</b>	
	<b>BeCl<sub>2</sub></b>
<b>magnesium nitride</b>	
	<b>K<sub>2</sub>O</b>
<b>aluminum oxide</b>	
	<b>BI<sub>3</sub></b>
<b>aluminum nitride</b>	
<b>iron(III) oxide</b>	
	<b>FeO</b>
<b>iron(III) chloride</b>	
	<b>CuCl</b>
<b>copper(II) chloride</b>	
	<b>PbO</b>
<b>lead(IV) bromide</b>	

Now use the “Binary ionic compounds” resource at the eChalk website to check your answers.